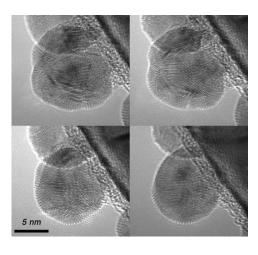
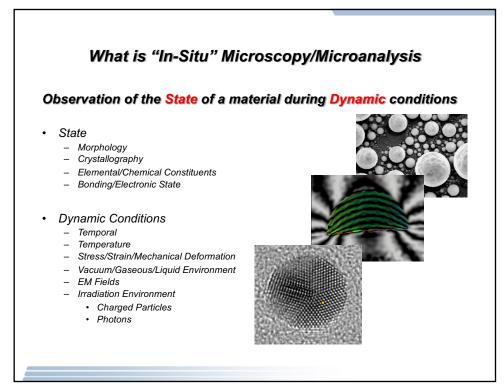
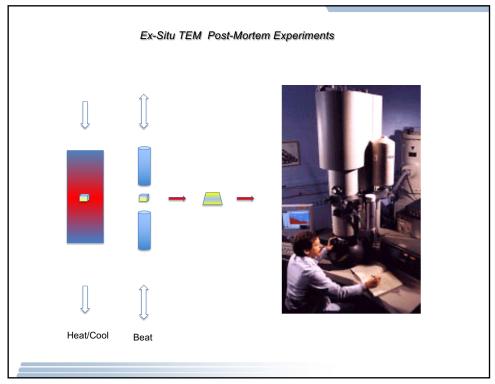
In-situ Transmission Electron Microscopy

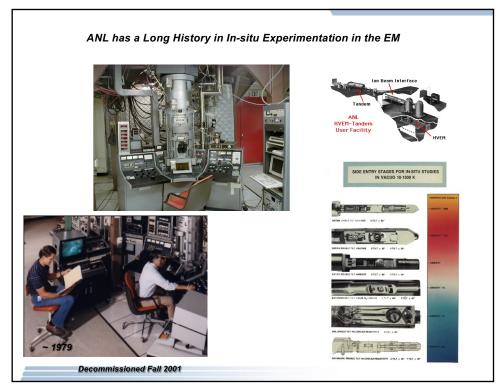


1

"In-Situ" Microscopy & Microanalysis Observation of the State of a material during Dynamic conditions • Dynamic Conditions - Temporal - Temperature - Stress/Strain/Mechanical Deformation - Vacuum/Gaseous/Liquid Environment - EM Fields - Irradiation Environment • Charged Particles • Photons • State - Morphology - Crystallography - Bonding/Electronic Structure - Elemental/Chemical Constituents







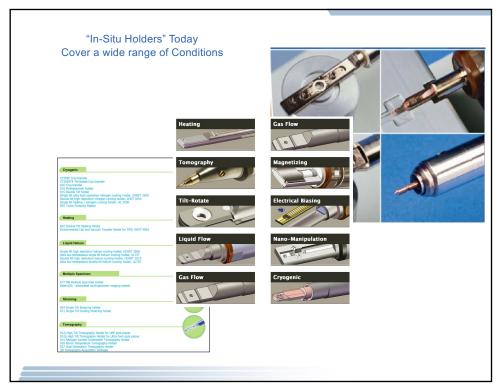
In-situ Characterization: The game is changing

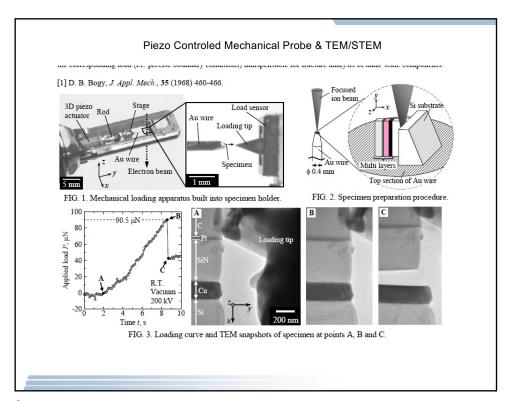
- •Improved Electron Optics
 - •Improved Performance
 - •Potentially Improved Experimental Space
- •In-Situ Holders/Environments
- •Improved Detector/Geometry

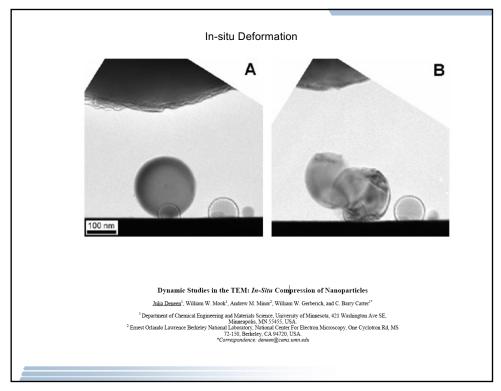
 - •Higher Resolution •Higher Speed DAQ •Higher Efficiency
- Computationally Mediated Experiments
 Exploit Electron Solid Interactions for
 State-of-the-Art Materials Characterization

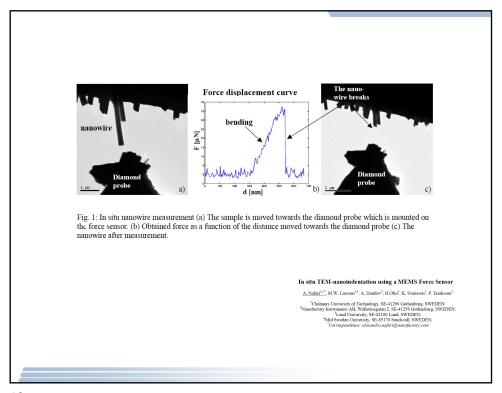
New Challenges:

- In-situ observation of growth processes at atomic resolution in growth environment
- Simultaneous imaging of hard/soft components
- Dynamics Fast detection schemes, detectors, and sources
- In-situ / high-spatial resolution elemental spectroscopy



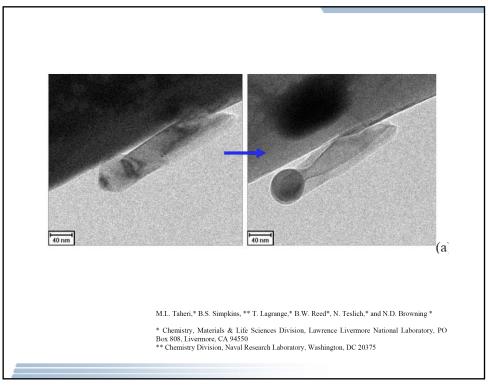


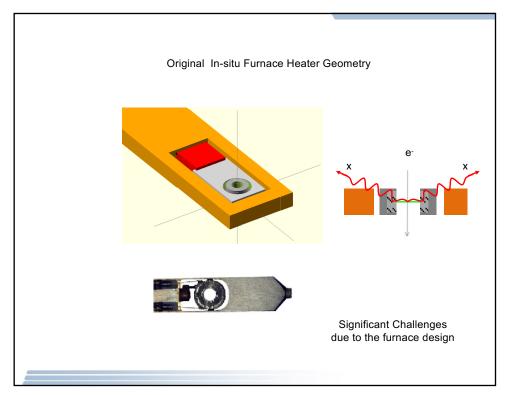


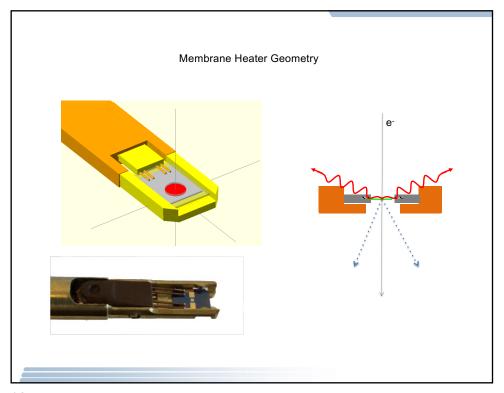


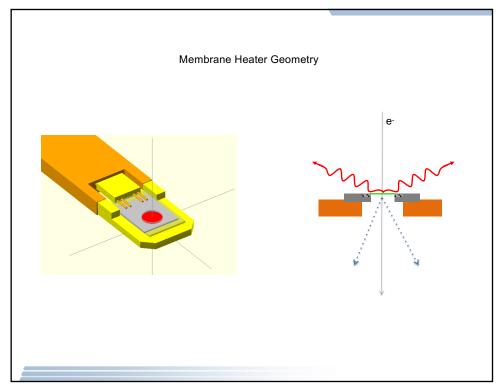
Gas Injection System heating element gas injection nozzle Fig.1 A gas injection-specimen heating holder developed for use with conventional TEMs. Environmental transmission electron microscopy using a conventional TEM and a gas injection-specimen heating holder T. Kamino*, T. Yaguchi*, A. Watabe*, H. Saka** and K. Kishita*** * Hitachi High-Technologies Cop., 11-11 Ishikaw-to, Hitachinaka, Iberaki, 312-0057 Japan **Nagoya University, Furn-cho, Chitachinaka, University, Furn-cho, Chitachinaka, Iberaki, 312-0057 Japan ***Material Analysis Dept. Toyota Motors Cop., 1 Toyota-cho, Toyota 471-8572 Japan

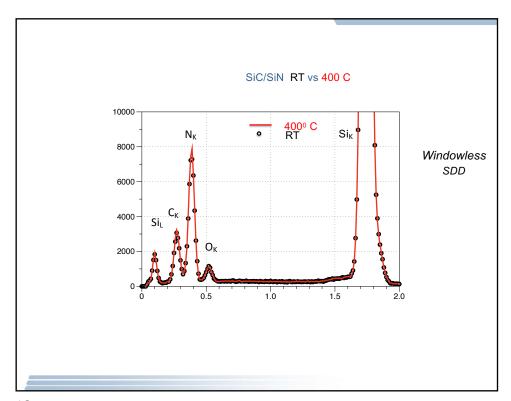
11

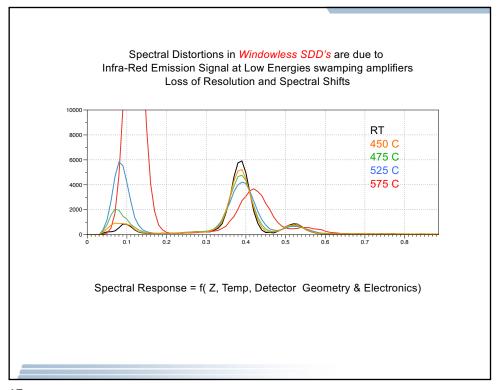


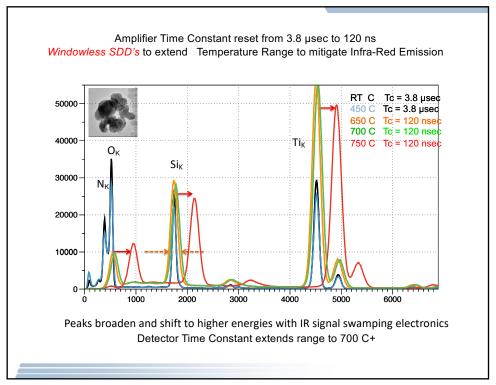


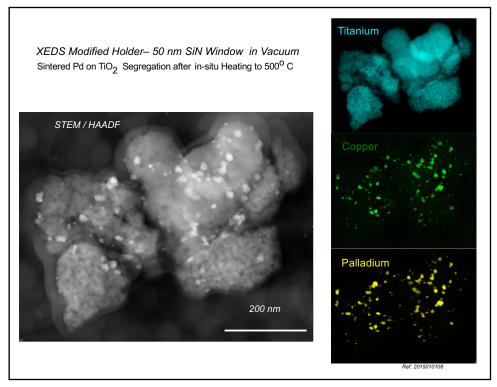


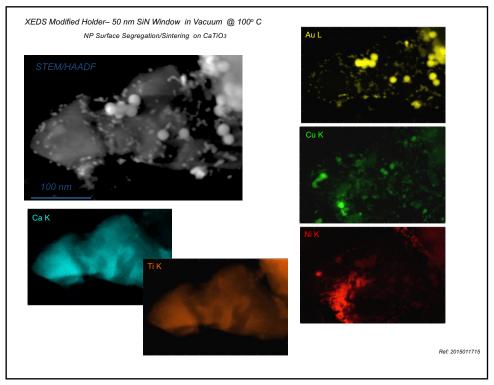


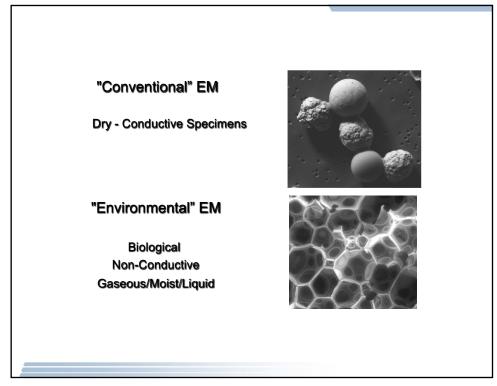


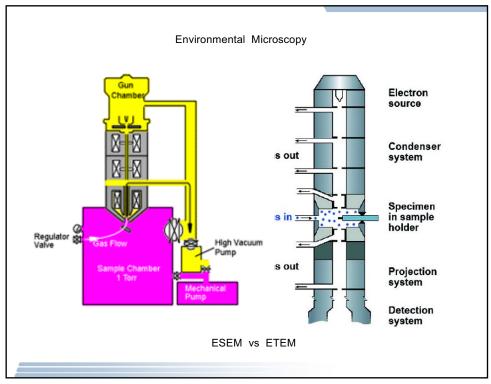


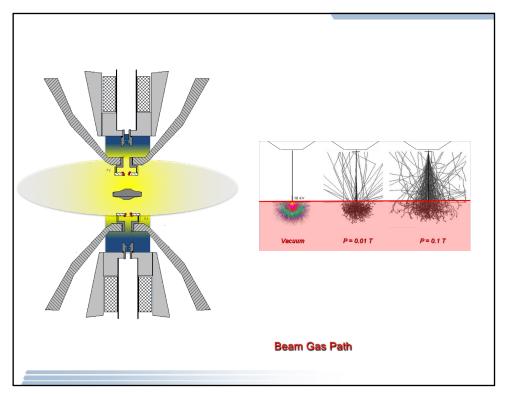


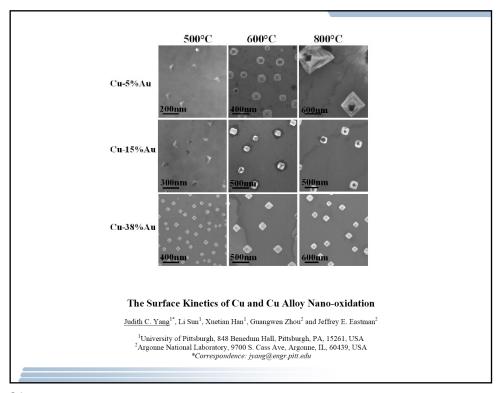


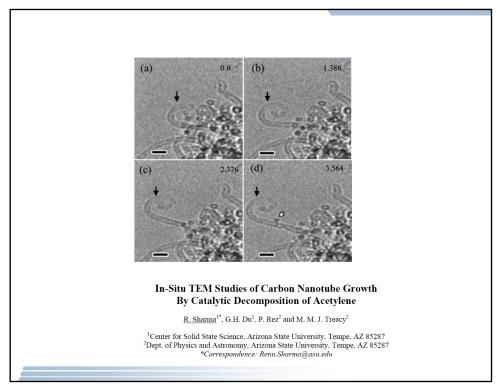


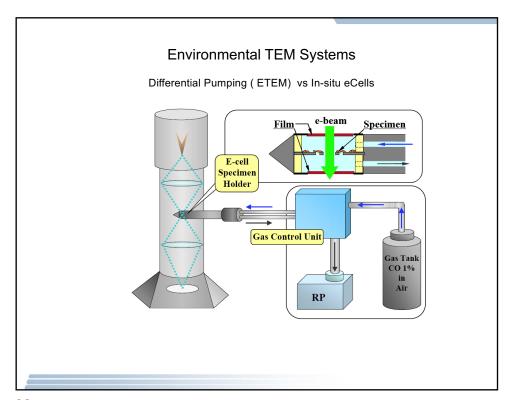


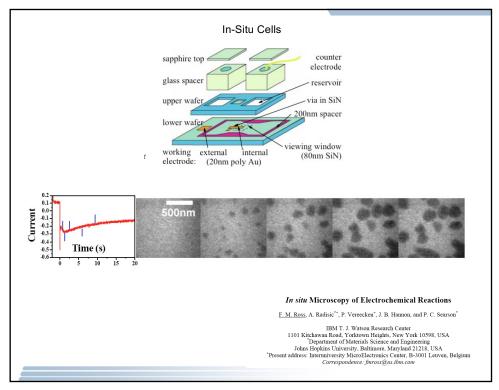


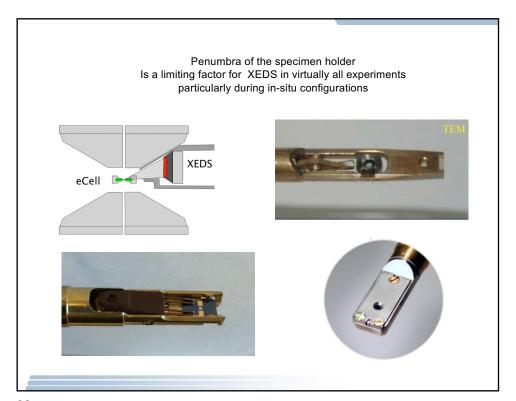


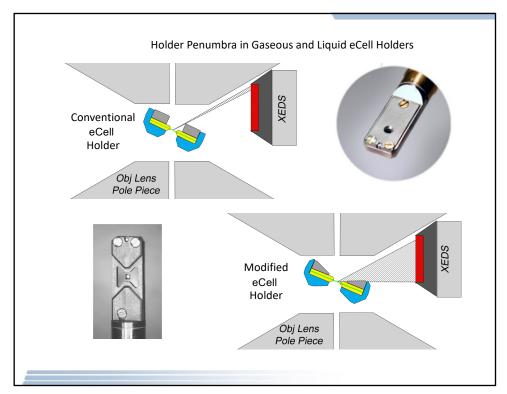


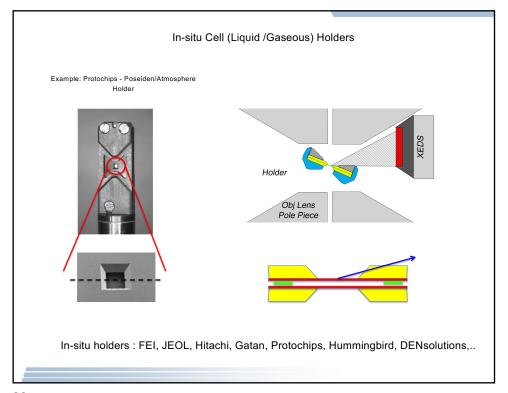


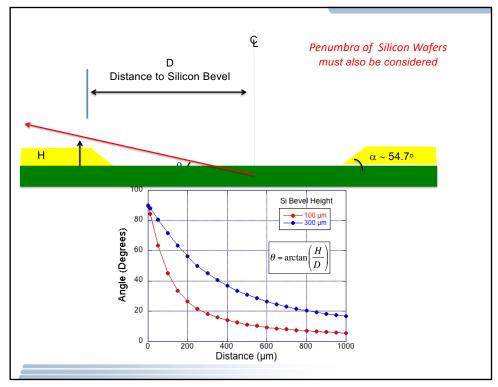


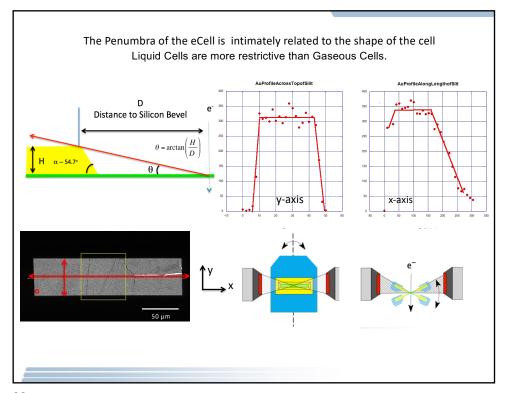


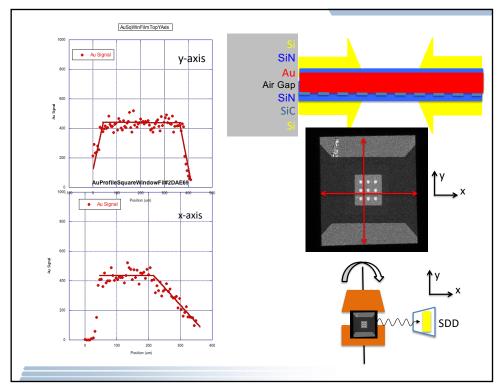


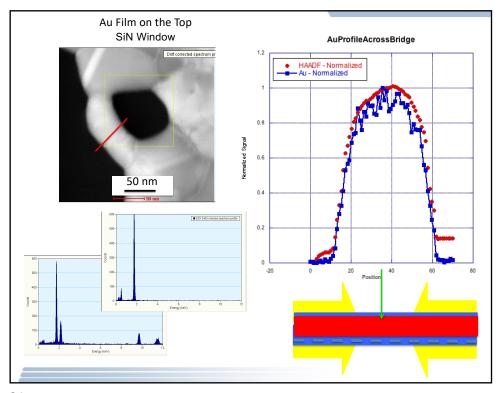


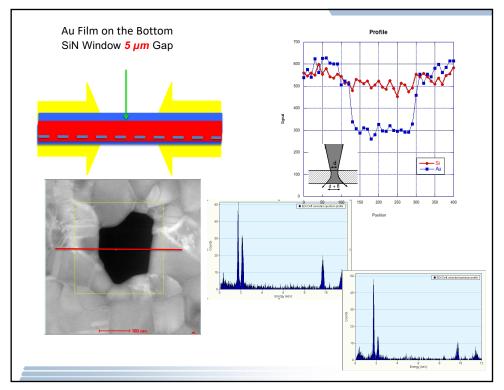


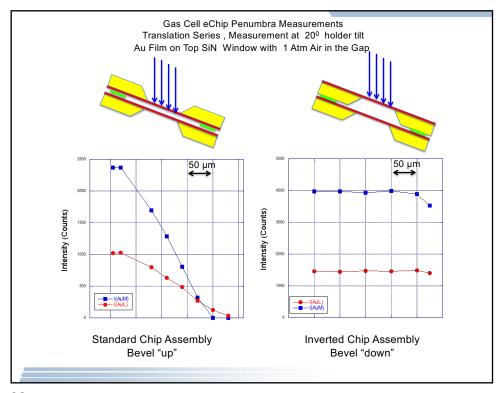


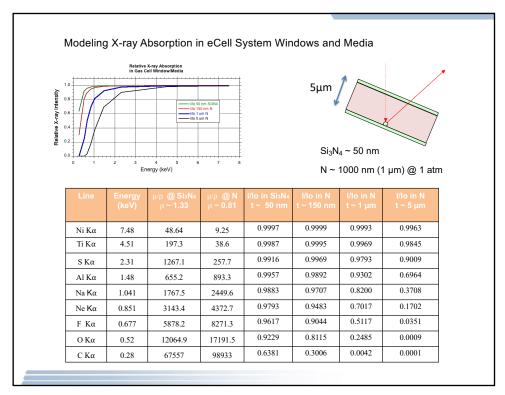


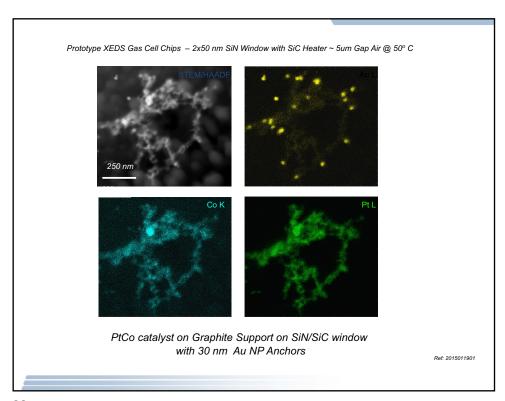


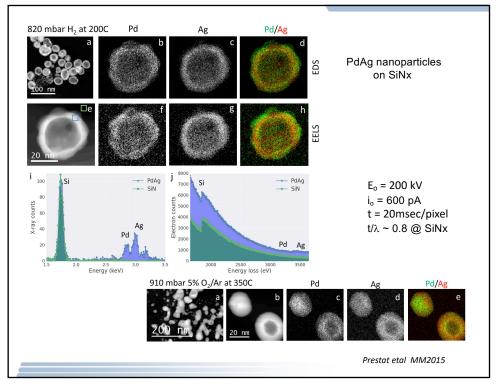


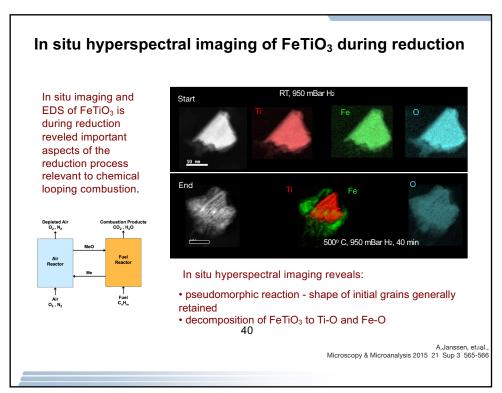










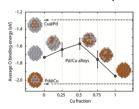


processes.

In situ hyperspectral imaging of PdCu/TiO₂ during H₂ reduction

Pd-M-TiO₂ catalysts (M = Sb, Bi, Sn, Ct version, Y-BA: benzyl acetate yield, Yields, S-BA: benzyl acetate selectivity.

Models of PdCu Catalysts



Tang etal. J. Phys. Chem. Lett. 2011, 2, 1328–1331

- Understanding the chemical changes during oxidation/reduction is important in many
- Supported bi-metal nanoparticle catalysts are key materials for diverse conversion processes throughout chemical industries.
- Bimetallic catalysts frequently offer improvements in activity, selectivity, and stability compared to their monometallic equivalents for diverse reactions such as catalytic reforming, hydrotreating, emissions controls, and biomass conversion.
- Evolution of catalyst microstructure and composition is of significant interest due to its role in understanding processing.

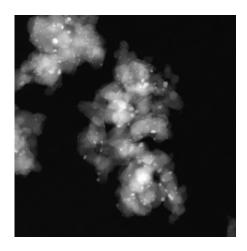
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Pd Cu on TiO₂

PdCu catalyst is representative of a large group of bimetallic catalysts useful for a range of chemical conversions including catalytic reforming, hydrotreating, emissions controls, and biomass conversion.

Studied at:

~ 1 ATM of Air & H₂ RT->550°C



M. A. Kulzick, P. J. Dietrich, E. Prestat, M. Smith, M. G. Burke, S.J. Haigh N.J. Zaluzec

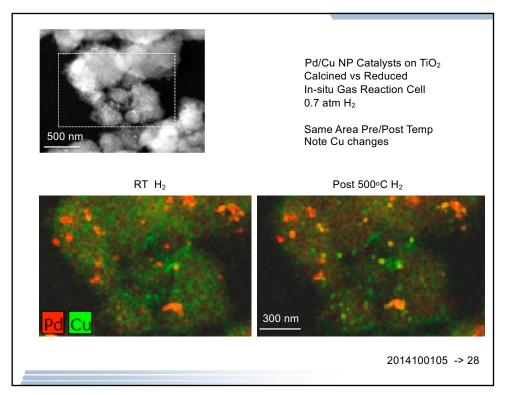


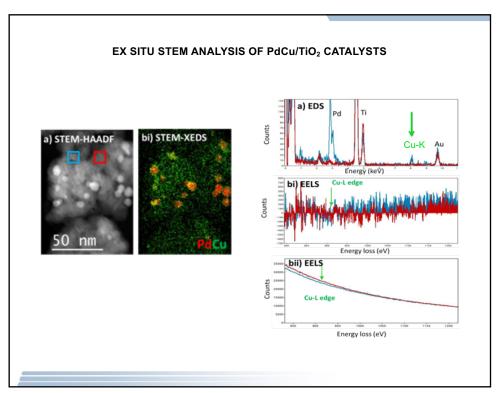


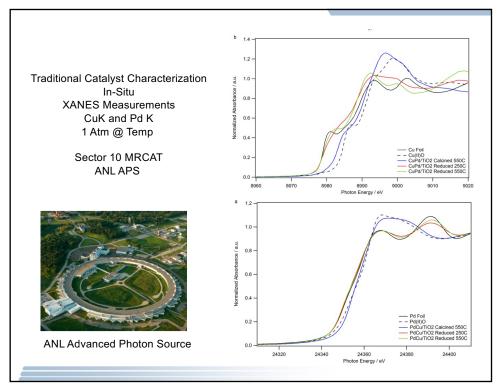


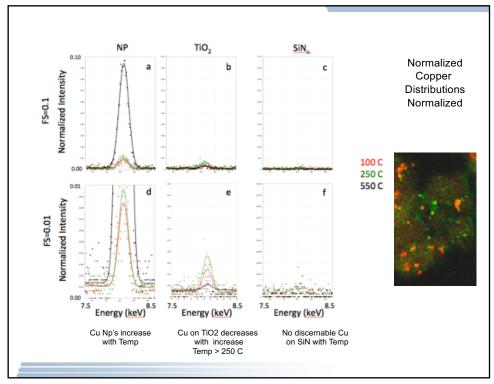


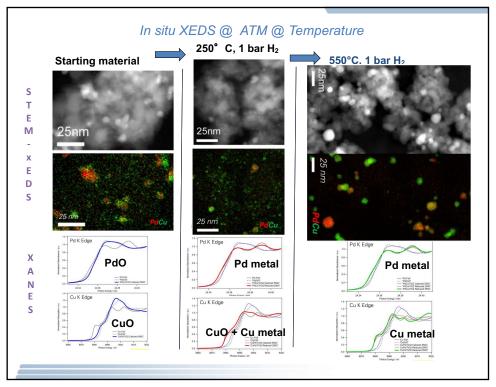


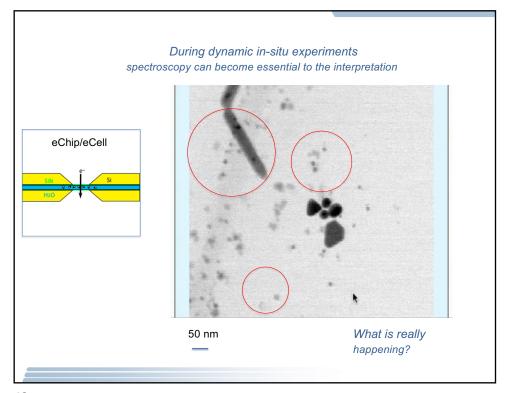


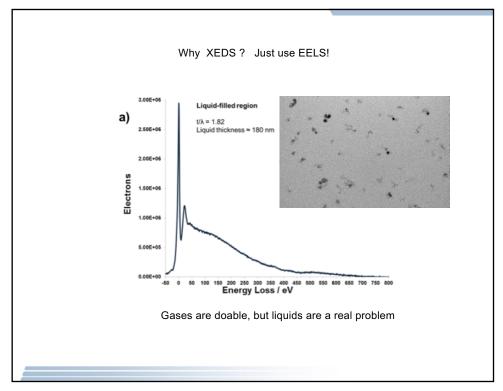


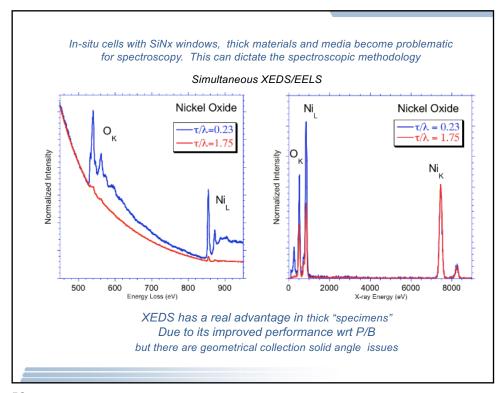


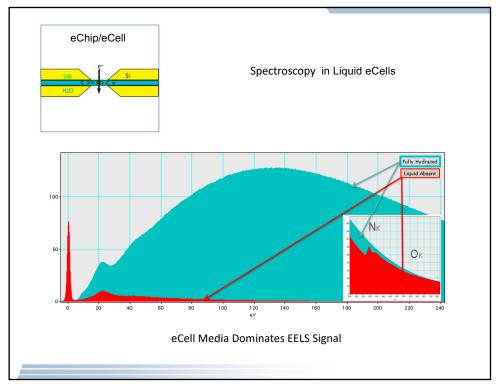


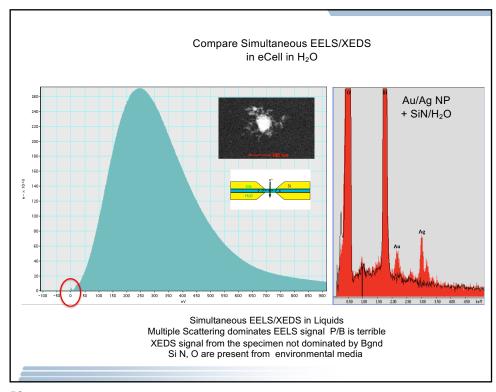


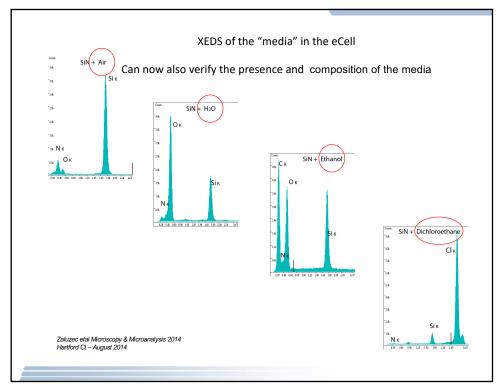


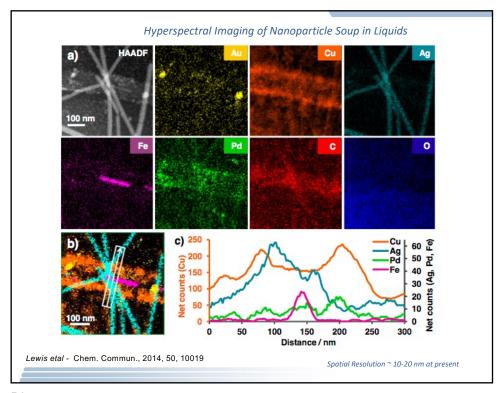


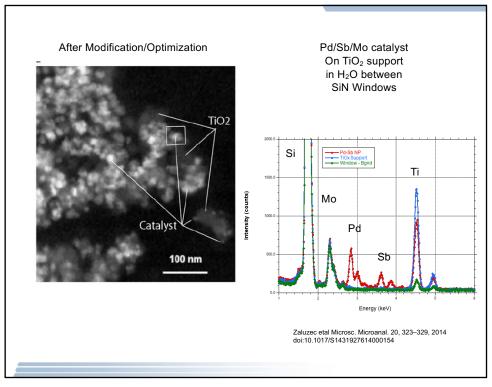


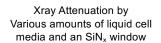


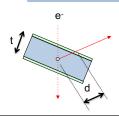












Energy (keV)	μ/ρ @ Si₃N₄ ρ ~ 1.33	μ/ρ @ H ₂ O ρ ~ 1	I/Io in Si₃N₄ d ~ 50 nm	I/lo in H₂O d ~ 150 nm	I/Io in H₂O d ~ 500 nm	I/lo in H₂O d ~ 1 µm
7.48	48.64	12.98	0.9997	0.9998	0.9994	0.9987
4.51	197.3	56.22	0.9987	0.9992	0.9972	0.9944
2.31	1267	390.6	0.9916	0.9942	0.9807	0.9617
1.48	655.2	1418	0.9957	0.9789	0.9315	0.8677
1.041	1767	3943	0.9883	0.9426	0.8211	0.6742
0.851	3143	7050	0.9793	0.8997	0.7029	0.4941
0.677	5878	13678	0.9617	0.8145	0.5046	0.2547
0.52	12064	29376	0.9229	0.6436	0.2302	0.0530
0.28	67557	176560	0.6381	0.0707	0.0001	0.00000001
	7.48 4.51 2.31 1.48 1.041 0.851 0.677	(keV) $\rho \sim 1.33$ 7.48 48.64 4.51 197.3 2.31 1267 1.48 655.2 1.041 1767 0.851 3143 0.677 5878 0.52 12064	(keV) p = 1.33 p = 1 7.48 48.64 12.98 4.51 197.3 56.22 2.31 1267 390.6 1.48 655.2 1418 1.041 1767 3943 0.851 3143 7050 0.677 5878 13678 0.52 12064 29376	(keV) $\rho \sim 1.33$ $\rho \sim 1$ $d \sim 50 \text{ nm}$ 7.48 48.64 12.98 0.9997 4.51 197.3 56.22 0.9987 2.31 1267 390.6 0.9916 1.48 655.2 1418 0.9957 1.041 1767 3943 0.9883 0.851 3143 7050 0.9793 0.677 5878 13678 0.9617 0.52 12064 29376 0.9229	(keV) $\rho \sim 1.33$ $\rho \sim 1$ $d \sim 50 \text{ nm}$ $d \sim 150 \text{ nm}$ 7.48 48.64 12.98 0.9997 0.9998 4.51 197.3 56.22 0.9987 0.9992 2.31 1267 390.6 0.9916 0.9942 1.48 655.2 1418 0.9957 0.9789 1.041 1767 3943 0.9883 0.9426 0.851 3143 7050 0.9793 0.8997 0.677 5878 13678 0.9617 0.8145 0.52 12064 29376 0.9229 0.6436	(keV) $\rho \sim 1.33$ $\rho \sim 1$ $d \sim 50 \text{ nm}$ $d \sim 150 \text{ nm}$ $d \sim 500 \text{ nm}$ 7.48 48.64 12.98 0.9997 0.9998 0.9994 4.51 197.3 56.22 0.9987 0.9992 0.9972 2.31 1267 390.6 0.9916 0.9942 0.9807 1.48 655.2 1418 0.9957 0.9789 0.9315 1.041 1767 3943 0.9883 0.9426 0.8211 0.851 3143 7050 0.9793 0.8997 0.7029 0.677 5878 13678 0.9617 0.8145 0.5046 0.52 12064 29376 0.9229 0.6436 0.2302

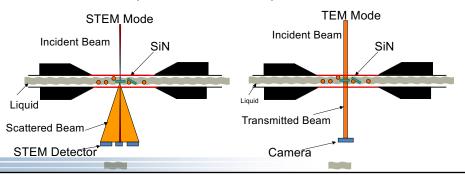
Calculation of x-ray absorption as a function of pathlength for a range of characteristic x-ray lines in SiN_x window and various amounts of liquid H_2O . As I/Io approaches unity the effects of x-ray absorption can be ignored.

Limitations of liquid environmental TEM studies

- · Electron scattering by liquid
- Sample preparation
- Mostly limited to materials such as nanoparticles or nanowires

"Bulk" metals/alloys in liquid environment studies?

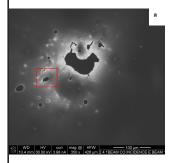
- Preparation of suitable metal specimens challenging
- Dimensional requirements for the liquid in situ eCell

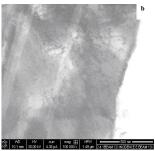


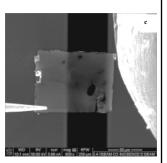
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Hybrid Specimen Preparation Methods to Study Conventional Materials

'Cut-and-Paste' approach using Dual Beam FIB





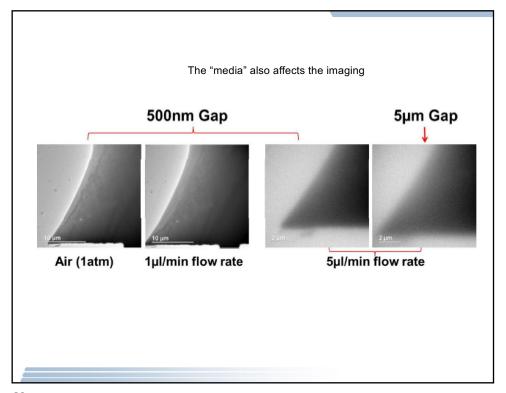


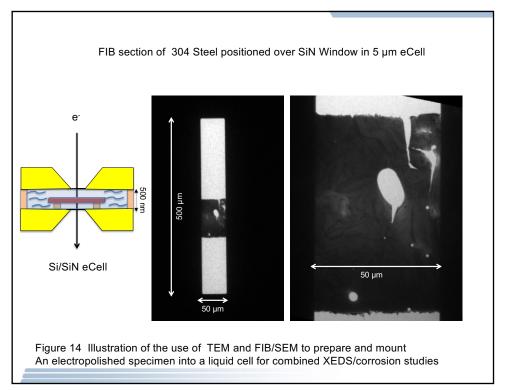
Quanta 3D DF-STEM image

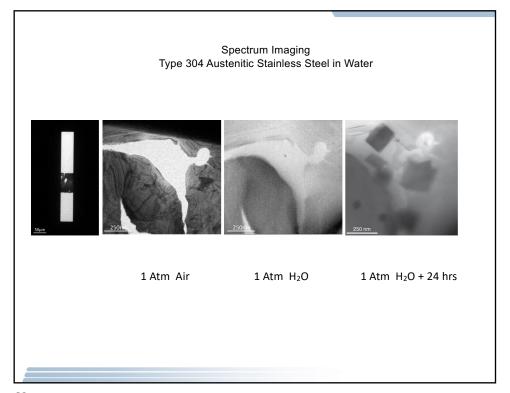
BF-STEM image

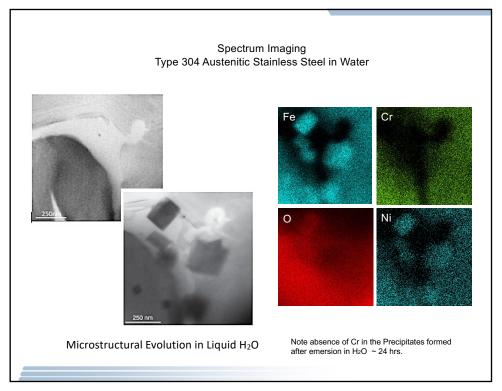
nage "cut" sample placed on TEM liquid cell window (50 μm wide)

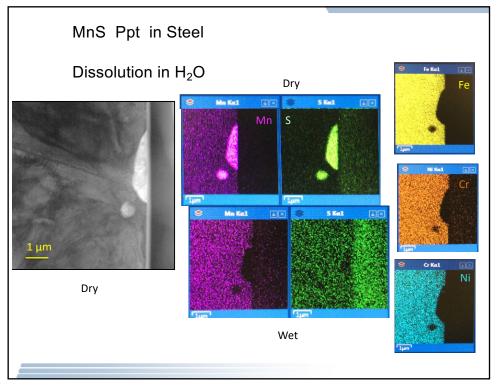


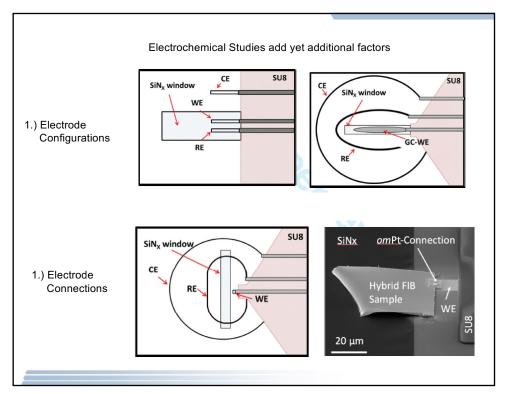


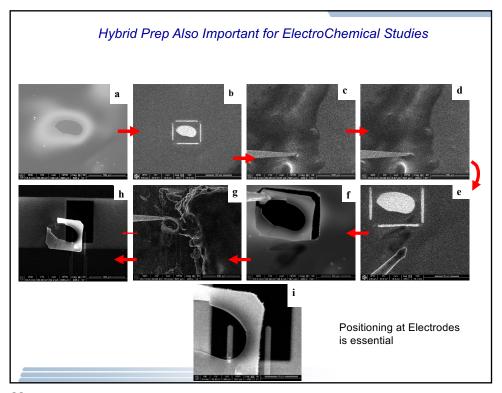


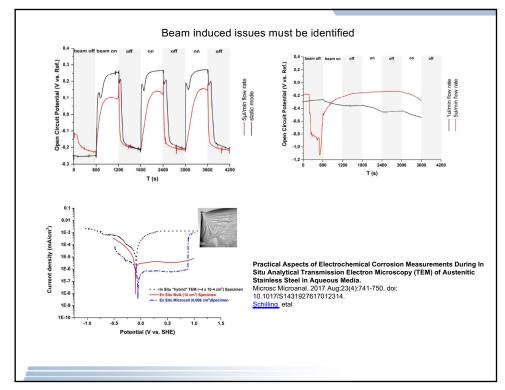












Hydrated electrons ??

In all microscopy experiments, the electron beam interacts with the sample resulting in the generation of radical and molecular species, which for water include e^h (hydrated electrons), OH, $H^\star,\,H_2,\,O_2,$ and $H_2O_2.$ The hydrated electrons, oxidizing agents, and gaseous species can cause, respectively, reduction and precipitation of cations from solution, dissolution of metals, and nucleation and growth of bubbles.

$$e^{-}$$
 (aq) + H_2O ——> H + OH^{-}
 e^{-} (aq) + H^{+} ——> H

The hydrated electron can act as a reducing agent. For example:

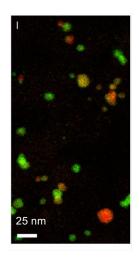
$$Cu^{2+} + e^{-} (aq) \longrightarrow Cu^{+}$$

 $[Cu^{2+} + H-----> Cu^{+} + H^{+}]$

Alot of work is still needed to understand and correctly measure and interpret all liquid TEM/STEM experiments.

Why was in situ Correlative Important?

- We can observe different details
 - Reveals nanometer scale concentration changes
 - But details not apparent without high resolution
- XANES invaluable but needed supplemental info
- XEDS saw all the critical elements
 - EELS is valuable, but not applicable here
- We can localize specific areas
 - In complex materials this is very important
 - Thermal drift adds complexity
- Beware beam induced changes!



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